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Easily accessible appendices provide the necessary mathematics, principles of electrical circuits, and basics of digital simulation. The breadth of coverage insures that this volume will be valuable not only to students in chemistry, biochemistry, industrial chemistry, chemical engineering, and materials science, but to researchers needing proper introduction to electrochemistry.

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INTRODUCTION 1.1 The scope of electrochemistry 1.2

The nature of electrode reactions 1.3 Thermodynamics and kinetics 1.4 Methods for studying electrode reactions 1.5 Applications of electrochemistry 1.6 Structure of the book 1.7 Electrochemical literature

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The first part introduces the fundamental principles of thermodynamics, kinetics, and mass transport associated with electrode reactions. The second part considers experimental methods that are available to study electrode and electrochemical processes, such as steady-state with forced convection, linear sweep, step-pulse voltametric techniques and impedance, modern surface analysis, and

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profiles of the various electrochemical methods.

Alternatively, the Nernst Einstein or Stokes Einstein relations discussed in Chapter 2 may be used to estimate values of D . The next question is: what is the variation of concentration with time due to diffusion? The variation is described by Fick's second law which, for a one-dimensional system, is $2c$ (5.2) $\times 2$ This is easily obtained

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Contents. Part 1 Principles: electrochemical cells - thermodynamic properties and electrode potentials-- the interfacial region-- fundamentals of kinetics and mechanism of electrode reactions-- mass transport-- kinetics and transport in electrode reactions.

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